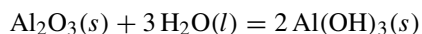


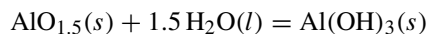
Chapter 2 Problemss

1. Calculate the volume change for the reaction



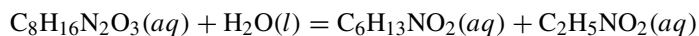
$\text{Al}(\text{OH})_3(s)$ is another way of writing the formula for gibbsite.

2. Calculate the volume change for the reaction



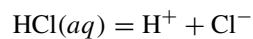
$\text{AlO}_{1.5}(s)$ is another way of writing the formula for corundum.

3. Calculate $\Delta_r V^\circ$ for reaction (2.3),



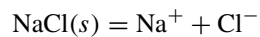
4. Why are the V° values for all the gases the same in Appendix B? Calculate this V° from data in Appendix A.

5. Calculate $\Delta_r V^\circ$ for the reaction



You should get zero. Why?

6. Calculate $\Delta_r V^\circ$ for the reaction



Note that the standard volume of many ions is negative. How can any substance have a negative volume?